

Express Mail No. EV355032575US

APPLICATION FOR UNITED STATES LETTERS PATENT

Applicants: Howard G. Ohlhausen
Jerome H. Ludwig

Title: A CLEANING AND MULTIFUNCTIONAL COATING COMPOSITION
CONTAINING AN ORGANOSILANE QUATERNARY COMPOUND
AND METHODS OF USING

**David J. Josephic
WOOD, HERRON & EVANS, L.L.P.
2700 Carew Tower
Cincinnati, OH 45202**

Att Dock #: REDVP-09

Specification

**A CLEANING AND MULTIFUNCTIONAL COATING COMPOSITION
CONTAINING AN ORGANOSILANE QUATERNARY COMPOUND AND
METHODS OF USING**

FIELD OF THE INVENTION

This invention relates to a cleaning composition which also provides a multifunctional coating on the clean surface to render it water and soil repellent, as well as antimicrobial. Organosilane quaternary
5 nitrogen compounds are formulated with hydrogen peroxide in aqueous media to provide the cleaning and multifunctional coating composition.

BACKGROUND OF THE INVENTION

Rain and coastal spray from lakes and oceans contain enough dissolved gases, carbonic acid, chlorides, sulfates, nitrates,
10 ammonia and particulate matter to make them chemically active. Similarly, most ground and municipally supplied water, while generally safe for drinking, cooking, washing and bathing, still contains enough

salts, hard water minerals, bacteria, organic and inorganic contaminants and water treatment chemicals to also be chemically active.

When either source of water is allowed to evaporate on most siliceous and nonsiliceous surfaces such as glass, porcelain, ceramic, marble, granite, polyurethane, polyester, polyacrylic, melamine/phenolic resins and polycarbonate, it combines with ordinary soil to promote the formation, adhesion and build-up of insoluble mineral deposits, lime scale, rust and soap scum. This ongoing exposure to chemically active water contributes directly to increased soiling and to the costly surface damaging cycle that results from scraping and scrubbing with harsh cleaners and cleansers to remove the insoluble buildup.

Thus, until rain, ground and municipally furnished water is completely deionized and demineralized, any kitchen and bath surface exposed to water and water-containing spills and splatters (such as sinks, counters, ranges, ovens, cooking and washing appliances, bathtubs and shower enclosures) remains subject to the adhesion and buildup of water insoluble soil.

To remove water insoluble soil, homemakers and cleaning professionals have turned to stronger and stronger cleaning aids including powerful acids and alkalis, penetrating solvents and a broad range of abrasive cleaners. The inevitable result of using such harsh cleaning aids is a surface destroying cycle. The stronger the cleaning action, the more it opens the microscopic pores and increases the hydrophilicity of the

surface. The more open and vulnerable the pores, the deeper the new soil penetrates and the greater its adhesion and build-up, and the stronger the cleaning action required to remove the soil with greater and greater destruction of the surface finish. Although microscopic at first, this cycle
5 ultimately leads to surfaces that are visibly rough, exhibit the loss of color and sheen, and become more prone to soiling.

It has been established that the soil and water repellency of surfaces can be improved by the use of silicone compositions. There has been extensive research and development involving silicone compositions
10 for rendering surfaces water repellent. For example, U.S. Pat. No. 2,612,458 (Stedman) discloses the application of substituted polysilicanes to windshields to achieve repellency. U.S. Pat. No. 2,923,653 (Matlin et al.) discloses improved compositions employing alkoxy groups in the polysilicane to improve the repellency. U.S. Pat.
15 No. 2,962,390 (Fain et al.) describes a paste containing a solid rubbing agent and an alkyl alkoxy silane which when rubbed on a glass surface provides repellency. U.S. Pat. No. 3,244,541 (Fain et al.) discloses acidic solutions of alkyl alkoxy silane monomers that produce rain repellent films on glass and which are also solvent resistant.

20 U.S. Pat. No. 3,579,540 (Ohlhausen) discloses water repellent film-forming compositions of alkylpolysiloxanes and acid, or alkylpolysiloxanes, acid and solvent which result in durable and effective water repellent films on various surfaces. Further improvements in

solventless compositions for treating porous and nonporous surfaces have been made as disclosed in U.S. Pat. No. 6, 432,181 (Ludwig and Ohlhausen). The '181 patent satisfied a need which existed for soil and water repellent compositions that eliminate solvents and utilize the silicone more effectively and economically. Additional improvements in soil and water repellent compositions were made by providing physiologically acceptable compositions that were nonirritating to the skin of the user as disclosed in U.S. Pat. Appln. Ser. No. 09/941,896, filed August 28, 2001 (Ludwig and Ohlhausen), now U.S. Pat. No. _____.

Many different types of hard and soft surfaces have also been rendered antimicrobial by coating with different agents. Organosilane quaternary nitrogen compounds have also been employed effectively in eliminating and/or reducing microbial contamination when applied to a variety of surfaces. For instance, bacterial, viral and fungal contamination may be eliminated or reduced when such organosilane quaternary compounds are applied to surfaces. Commercially available quaternary ammonium organosilanes which have been used for this purpose include 3-(trimethoxysilyl)propyldimethyloctadecyl ammonium chloride, 3-(trimethoxysilyl)propyldidecylmethyl ammonium chloride, and 3-(trimethoxysilyl)propyltetradecyldimethyl ammonium chloride.

Notwithstanding the improvements that have been made over many decades of research and development regarding cleaning or coating

compositions for various surfaces, there still exists the need for further improvements.

SUMMARY OF THE INVENTION

This invention relates to a composition for cleaning a surface
5 and providing it with a bonded multifunctional coating. The cleaning and
coating composition comprises a cationic organosilane quaternary
ammonium compound (sometimes hereinafter simply "organosilane quat")
and hydrogen peroxide which is bondable onto said surface. The
components are used in effective amounts in aqueous media for cleaning
10 a surface and for bonding a multifunctional coating onto said surface
thereby rendering it (a) water and soil repellent and (b) antimicrobial.

It has been found that the inventive coating composition has
surfactant properties which enable a soiled surface to be cleaned and
simultaneously rendered water and soil repellent, as well as antimicrobial.
15 Accordingly, everyday surfaces that are soiled with everyday household
soil that result from cooking, eating, washing, etc. may be simultaneously
cleaned and provided with a bonded coating that rather permanently repels
water and soil while having antimicrobial properties.

It has also been discovered that the organosilane quat in
20 combination with the hydrogen peroxide provides synergistic results upon
bonding onto surfaces. The inventive composition containing the
combination of the organosilane quat and hydrogen peroxide components
provides unexpectedly improved bonding and durability of the composition

on various surfaces. In other words, the results achieved with the combined components in the composition exceeds the expected algebraic sum of the activity of each component when separately used on the surface. These synergistic activities contribute to the unique compositions and methods of this invention.

In general, the organosilane quat is used in the composition in an amount up to about 3% by weight and the hydrogen peroxide is used in an amount up to about 8% by weight of the composition. More preferably, the organosilane quat is used in an amount of up to about 1% by weight and hydrogen peroxide is used in an amount of about 3 to about 6% by weight.

The coating composition may further contain a solvent such as an alcohol, polyol, glycolether and mixtures thereof, for example, glycol, propylene glycol monomethyl ether, methanol, ethanol or isopropanol. Further, preferably the aqueous media is acidic, wherein the pH is on the order of about 2 to about 5. The composition is also preferably formulated with deionized water.

The invention also involves a method for treating everyday surfaces as found in homes, offices, vehicles, etc. with a composition that provides a clean surface and a multifunctional coating. For example, everyday surfaces that are soiled with spills, splatters and blemishes that results from cooking, eating, washing, etc. may be cleaned and provided with the multifunctional coating upon application of the inventive

composition. For example, a liquid inventive composition can be applied by wiping or spraying onto a soiled surface in an effective amount and the soil is removed by wiping. Upon such application, the surface becomes clean and a multifunctional coating is bonded onto the surface thereby forming a clean, soil and water repellent and antimicrobial surface.

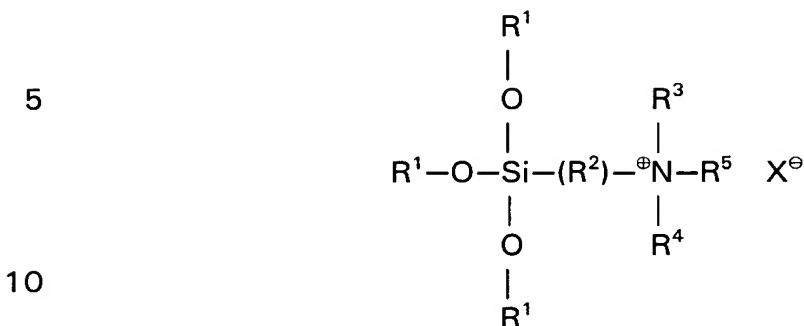
A further understanding of the invention, its various embodiments and operating parameters will be apparent with reference to the following detailed description.

DETAILED DESCRIPTION OF THE INVENTION

In accordance with the above summary, the objectives of this invention are to provide compositions and methods for simultaneously cleaning, disinfecting and rendering everyday surfaces water and soil repellent in essentially one step. The invention provides compositions that are (1) VOC compliant and biodegradable, (2) high-performance, penetrating cleaners, (3) broadband disinfectants, (4) bondable, durable water and soil repellents that are resistant to removal by soaps, solvents, detergents and mild abrasives, (5) effective on an extraordinary range of everyday household and vehicular surfaces, (6) storage stable and economical in terms of performance and coverage, and (7) that can be used or applied with ordinary spray-and-wipe techniques.

In accordance with the best mode of this invention, the multifunctional coating composition includes a cationic organosilane

quaternary ammonium compound which is bondable onto said surface as defined by the formula:



wherein R¹=hydrogen and/or C₁ to C₄ alkyl; R²=divalent hydrocarbon radical with C₁ to C₈ carbon atoms, R³=hydrogen or C₁ to C₄ alkyl, R⁴=hydrogen or C₁ to C₁₀ alkyl, R⁵=C₁₀ to C₂₂ saturated or unsaturated hydrocarbon radical and X=halide (preferably chloride or bromide), carboxylate (acetate, glycolate), sulfonate, hydroxide, sulfate, or phosphate.

The organosilane quat is contained in an amount of up to about 1 to 3%, preferably about 0.4-0.7% and the hydrogen peroxide is in an amount up to about 8%, preferably about 3 to 6%, in an acidic deionized aqueous media. The pH of the acidic media is about 2 to about 5, preferably about 3. The composition may further contain a solvent selected from the group of an alcohol, polyol, glycolether and mixtures thereof, as mentioned above, more preferably, methanol, ethanol or isopropanol.

The cationic organosilane quaternary ammonium compound is selected from the group consisting of

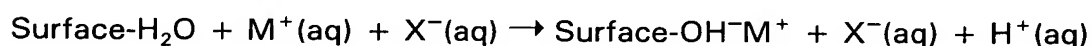
3-(trimethoxysilyl)propyldimethyloctadecyl ammonium chloride,
3-(trimethoxysilyl)propyldidecylmethyl ammonium chloride,
3-(trimethoxysilyl)propyltetradecyldimethyl ammonium chloride,
3-(trimethoxysilyl)propyldimethylsoya ammonium chloride,
5 3-(trimethoxysilyl)propyldimethyloleyl ammonium chloride,
3-(trimethoxysilyl)propyloctadecyl ammonium chloride,
3-(trimethoxysilyl)propyloleyl ammonium chloride, and
3-(trihydroxysilyl)propyldimethyloctadecyl ammonium chloride. Other
suitable organosilane quats may be used, providing they function to clean
10 and bond a multifunctional coating onto a surface.

It has been found that the preferred compositions of this
invention contain organosilane quaternary ammonium compounds that
have surfactant activities which facilitate the cleaning of the surface and
also provide reactive groups for bonding with the surface to obtain the
15 desired cleaning and multifunctional coating capabilities. Accordingly, it
is preferred for the organosilane to have a hydrocarbon group such as a
 C_{10} - C_{22} saturated or unsaturated hydrocarbon group which facilitates the
surfactant, water repellent and antimicrobial activities.

Suitable quaternary organosilanes are described in further
20 detail with reference to the following Examples. As stated above, the
relative amounts of the quaternary silanes and hydrogen peroxide to
achieve the objectives of this invention, in general, are up to about 3%
and 8% by weight, respectively. More specific amounts for the most

preferred organosilane quats, like 3(trimethoxysilyl)propyldimethyloctadecyl ammonium chloride, are about 0.4 to about 0.7% by weight and about 3% by weight of hydrogen peroxide.

Quaternary ammonium compounds are known to bond to anionic surfaces by interaction with absorbed water on the surface. This can be described by the following equation.



wherein M^+ is an organic cation other than hydrogen ion and a hydrogen ion is displaced from the adsorbed surface water to become hydrated and to enter the aqueous phase as $\text{H}^+(\text{aq})$. Thus the organic cation (M^+) is "hydrogen bonded" to the anionic surface.

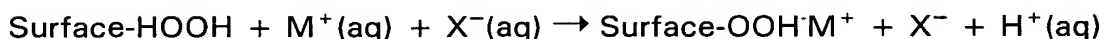
While not desiring to be bound by theory, it is believed helpful for a further understanding of the invention to postulate a mechanism by which the cationic organosilane quaternary compound is bonded to the surface in the presence of hydrogen peroxide. Hydrogen peroxide is well known for its antimicrobial activity and its oxidizing power which has been used to destroy a variety of toxic pollutants. In aqueous solution it is a stronger acid than water with reference to the following equation:



However, it has not been known before this invention to combine hydrogen peroxide with an organosilane quat to obtain the

cleaning and coating benefits described herein, especially the synergistic activities.

In order to understand the synergistic activities that have been demonstrated in the following Examples of this invention, it is
5 believed that hydrogen peroxide, which is very similar to water, will be adsorbed onto anionic surfaces in a similar manner as demonstrated by the following equation:



wherein M^+ is an organic cation other than H^+ , and H^+ is displaced from
10 adsorbed surface hydrogen peroxide to become hydrated and to enter the aqueous phase as $\text{H}^+(\text{aq})$. It is believed that the organic cation (M^+) is thus hydrogen bonded to the anionic surface associated with the hydroperoxide ion ($-\text{OOH}^-$) even stronger than when associated with the hydroxide ion ($-\text{OH}^-$) when water is on the anionic surface.

15 If the organic cation also contains a silane function ($-\text{Si-OR}$) which can hydrolyze to a silanol ($-\text{Si-O-H}$), additional bonding of a chemical nature can occur with the substrate silanols on the substrate surfaces or to another organic cation silanol which would lead to polymerization and crosslinking of the organo quaternary compounds on
20 the substrate surface. This chemical bonding leads to a durable coating which is not easily removed from the substrate surface.

With the additional cleaning capability of hydrogen peroxide on anionic surfaces and its ability to strengthen the hydrogen bonding of

the silane quaternary cation (M^+) to the anionic surface, thus facilitating the enhanced chemical bonding of the silane to the surface and subsequent crosslinking, an improved surface bonding or durability of the coating results from application of the compositions of this invention. It has been found that the combination of the cationic organosilane quaternary compound with hydrogen peroxide provides synergistic results. In other words, the resultant bonding and durability of this combination of components unexpectedly exceeds the summation of the individual components' activities, as demonstrated by the Examples which follow.

The invention may be more readily understood by the following detailed disclosure of preferred embodiments of the invention. It is to be understood that the terminology used herein is for the purpose of describing particular embodiments and is not intended to be limiting.

The term "abrasion resistant" as used herein refers to a surface, surface coating or finish that is resistant to damage or removal by washing, scraping or scrubbing with a mildly abrasive substance or process without visibly damaging to the surface or finish, as in scratching or blemishing the surface.

The term "antimicrobial" is used herein to describe the elimination, reduction and/or inhibition of microorganism growth such as mold, virus, fungus or bacteria.

The term "application", "applying" or "applied" as used herein means the treatment of a surface generally and (a) in the case of a liquid, by spraying and wiping, or agitating and wiping, or by wiping and polishing, (b) in the case of a cream, by wiping and/or polishing and (c) 5 in the case of a gel, by wiping and/or polishing. The application of the inventive composition is usually by spraying and wiping, wiping and polishing or wiping, agitating and polishing, although these forms of treatment are by exemplification and not restrictive.

As used herein, the term "bond", "bonded" or "bondable" 10 means the ability to strongly adhere the composition to the surface, as in the ability to bond a water & soil repellent finish, coating or characteristic to an otherwise water and soil accepting surface. As used herein, the composition is deemed "bonded" or "bondable" when it is resistant to removal by soaps, solvents, detergents or abrasive-type 15 cleansers that would not otherwise stain, blemish or damage an untreated form of the same surface.

As used herein, "multifunctional" means the process of achieving two or more discernable results from a single application, as in simultaneously or sequentially cleaning and coating a surface whereby the 20 coating also performs the function(s) of rendering the surface water repellent, soil repellent and/or antimicrobial.

The term "durable" or "durability" means long-lasting and not easily removed by washing and/or wiping using plain (tap) water,

soap solutions, detergent solutions, household solvents, mildly abrasive (non-damaging) cleansers or conventional cleaner/degreasers.

By the term "everyday household soil" is meant the spills, splatters and blemishes on a surface that result from cooking, eating,
5 drinking, washing, bathing and showering such as milk, coffee, tea, juices, sauces, gravies, food boil over, soap scum, water spots, mineral deposits, etc.

By the term "everyday vehicular soil" is meant the spills, splatters and blemishes on the exterior of a vehicular surface that result
10 from rain, sleet, snow, insects, mud and road grime, and on the interior of a vehicular surface that result from fingerprints, food spillage, plasticizer leaching, smoking, use of hair and deodorizing sprays and air circulation.

By the term "surface(s)" is meant the full range of hard or
15 soft surfaces, rather porous or non-porous, siliceous or non-siliceous, as exemplified by everyday surfaces and such as those used in the following examples which illustrate the compositions and methods of this invention. Examples of such surfaces include, without limitation metal, glass, plastics, rubber, porcelain, ceramic, marble, granite, cement, tile, sand,
20 silica, enameled appliances, polyurethane, polyester, polyacrylic, melamine/phenolic resins, polycarbonate, siliceous, painted surfaces, wood, and the like.

By the term "everyday surfaces" as used herein is meant the full range of surfaces in homes, offices, factories, public building and facilities, vehicles, aircraft and ships, and the like.

By the term "resistant to removal" is meant a coating or
5 surface finish that is not easily removed by washing or cleaning with conventional soaps, solvents, detergents, mildly abrasive cleansers or cleaner/degreasers that would not otherwise etch or damage an untreated surface of the same composition and construction.

As used herein, the term "soil repellent" means a surface
10 that exhibits reduced adhesion to, and buildup of, for example, everyday household and vehicular soil both before and after evaporation of the water component.

The terms "water repellent" and "water repellency" as used herein describe the hydrophobic nature or characteristic of a surface and
15 its ability to repel water as measured by the contact angle of a drop or droplet of distilled water on the surface. (Contact angles measured with rainwater, municipally furnished tap water or ground water are typically more variable and non-reproducible, and commonly measure up to 10° less than those using distilled or deionized water.) Generally, the
20 hydrophobicity of a discrete surface is rated in terms of its contact angle to water drops as follows:

Excellent Compact drops, well rounded, with bright sparkles
measuring 95° or more

	Good	Less rounded drops, but bright sparkles that exhibit slight spread, measuring 85° to 95°
	Fair	Visible flattening of the water drops, measuring 70° to 85°
5	Poor	Relatively flat water drops, exhibiting more spread of the water and measuring 50° to 70°.

EXAMPLES 1 - 22 - COMPARATIVE BONDING TESTS ON SURFACES

These tests utilize the degree and durability of the water repellency of the treated surface as a practical measure of bonding after application of the composition to the surface followed by scrubbing of the surface with an abrasive cleaner that is capable of removing the repellent coating from the surface without scratching. The degree of repellency was determined by the contact angle and appearance of water drops when applied to the scrubbed area on the surface. Coatings resulting from the application of (1) the inventive compositions of the organosilane quat and hydrogen peroxide were compared to coatings of (2) the organosilane quat, alone and (3) hydrogen peroxide, alone. Comparative coatings of other organic cationic species, with and without hydrogen peroxide, were also tested.

A. PREPARATION OF THE SURFACE

New 12"x12" glass mirrors were cleaned and made hydrophilic by scrubbing the surface with the soft cleanser MIRACLE SCRUB, a non-scratching, mildly abrasive hard surface cleanser available

from Unelko Corporation, Scottsdale, Arizona, and using a moist cellulose sponge. After cleaning, the mirror surface was rinsed with hot water to remove the MIRACLE SCRUB residue followed by a rinse with deionized water followed by drying with a paper towel. The cleaned mirrors were
5 then allowed to air dry for at least 24 hours. A surface is judged clean and free of surface soil when water spreads out and wets the surface.

B. PREPARATION OF TEST COMPOSITIONS

The test solutions for the Comparative Bonding Tests were all prepared according to the following formulations with and without
10 hydrogen peroxide.

A. Without Hydrogen Peroxide

6% (wt.) Isopropyl alcohol
0.4% (wt.) Organic Cationic
q.s. Deionized Water

B, With Hydrogen Peroxide

6% (wt.) Isopropyl alcohol
0.4% (wt.) Organic Cationic
3.0% (wt.) Hydrogen Peroxide
q.s. Deionized Water

15 In all compositions the pH was adjusted to 3 with 1.0 N hydrochloric acid.

C. COMPARATIVE BONDING TEST PROCEDURE

Two compositions formulated with the same organic cationic species, but one with and one without hydrogen peroxide were applied
20 to the same 12"x12" mirror. One half of the mirror was masked with paper and the other half was sprayed with the peroxide-free composition followed by spreading with a paper towel and wiping the surface until dry. The covering was reversed, and the other half of the mirror was

then treated in the same manner with the composition containing hydrogen peroxide.

The mirror was then rinsed with tap water followed by deionized water to remove any excess of the compositions from the surface. The rinsed mirror was then dried with a paper towel.

A small but equal amount of MIRACLE SCRUB was applied to a section of each half of the now-treated mirror surface. A moist cellulose sponge was used to scrub the surface of each test half using a circular motion and a moderate pressure on the sponge. Ten circular motions were used on each area. The MIRACLE SCRUB residue was then rinsed with tap water followed by a deionized water rinse. The mirror surfaces were then dried with a paper towel. The dried, scrubbed areas were then compared and evaluated for water repellency to determine the relative durability (degree of bonding) of the two surface coatings.

With the mirror in a horizontal position, several equal sized drops of deionized water were placed on the scrubbed areas of each section of the mirror surface and evaluated for repellency. The contact angle of the drops as well as the spread of the drops were assessed and rated in accordance with the standard for water repellency described in the definition of terms: Excellent - Good - Fair - Poor.

D. TEST RESULTS

Table 1 represents the comparative bonding test results on six organosilane quaternary compounds formulated according to the

above formulas A and B. In all tests, improved bonding or durability of the coating was observed in formulations containing hydrogen peroxide. Examples 1, 2, 3 and 6 are tetra-alkyl quaternary compounds. Examples 4 and 5 are dialkyl quaternary compounds. Examples 1, 2, 3, 4 and 5 are trialkoxysilanes. Example 6 is a trihydroxysilane.

TABLE 1
Organosilane Quaternary Compounds
Comparative Bonding Test Results

Example	Organosilane Quaternary Compound	Without H ₂ O ₂	With H ₂ O ₂
1	3(TMS)P*-dimethyloctadecyl-ammonium chloride	Good	Excellent
2	3(TMS)P*-dimethylsoya-ammonium chloride	Fair	Good-Excellent
3	3(TMS)P*-dimethyloleyl-ammonium chloride	Good	Good-Excellent
4	3(TMS)P*-octadecyl-ammonium chloride	Good	Excellent
5	3(TMS)P*-oleyl-ammonium chloride	Fair	Good
6	3-(trihydroxysilyl)propyldimethyloctadecyl ammonium chloride	Poor-Fair	Fair-Good

*3(TMS)P = 3(trimethoxysilyl)propyl

Table 2 represents the Comparative Bonding Test Results on two trialkyl quaternary compounds (Examples 7 and 8) and one tetraalkyl quaternary (Example 9) along with three trihydroxy silanes (Examples 11, 12 and 13). Examples 12 and 13 were employed at a 1 % active level in ethanol solution. All Examples with or without hydrogen peroxide were judged as "Poor" on the Comparative Bonding Test. This indicates that all of the silane and quaternary derived coatings were easily removed with a mild abrasive cleaner. This also indicates that there were no chemical bonds between the "coating" and the substrate. Perhaps only hydrogen bonding, which is a much weaker bond than a chemical bond, was holding the "coating" to the substrate.

It was also noted that prior to scrubbing of the treated surface with MIRACLE SCRUB, all treated surfaces were rated by the water drop test as follows:

15	Example 7:	Excellent - both with and without peroxide
	Example 8:	Good without peroxide Fair with peroxide
	Example 9:	Good - both with and without peroxide
	Example 10:	Fair - both with and without peroxide
	Example 11:	Poor-Fair - both with and without peroxide
20	Example 12:	Good - both with and without peroxide
	Example 13:	Good - both with and without peroxide

TABLE 2

Silane or Quaternary Compounds

Comparative Bonding Test Results

	Example	Silane or Quaternary Compound	Without H ₂ O ₂	With H ₂ O ₂
5	7	Octadecyldimethyl ammonium chloride	Poor	Poor
	8	Di(C ₈ -C ₁₀) alkyldimethyl ammonium chloride	Poor	Poor
	9	Alkyl (C ₁₄ , C ₁₆ , C ₁₂) dimethyl-benzyl-ammonium chloride	Poor	Poor
	10	Decyldimethyl ammonium chloride	Poor	Poor
	11	Chloropropyltrihydroxy silane	Poor	Poor
10	12	Trifluoropropyl-trihydroxy silane	Poor	Poor
	13	Tridecafluorooctyl-trihydroxy silane	Poor	Poor

To further evaluate the invention, additional kitchen and bath surfaces were treated with one of the preferred organosilane quaternaries of Example 1, i.e., 3-(trimethoxysilyl)propyldimethyloctadecyl ammonium chloride was used in all Examples 14-22-Table 3, employing the test formulations without and with hydrogen peroxide, as defined earlier, and with 3% hydrogen peroxide alone. Table 3, Kitchen and Bath Surfaces, Comparative Bonding Test Results, reports the evaluation on 8 additional everyday surfaces found in the kitchen and bath areas. The hard, glass-

like surfaces of Examples 14, 15, 16 and 22 were cleaned and treated with the organosilane quaternary compositions (with and without hydrogen peroxide) according to the A. PREPARATION OF THE GLASS TEST SURFACE previously described. The more porous surfaces of Examples 17-21 were cleaned with MIRACLE SCRUB in the same manner, however, there were 3 treatments of the surfaces with the compositions, employing a tap water rinse followed by deionized water rinse between each application and after the last application. The Comparative Bonding Test was then run on all the surfaces. The results are summarized in Table 3.

TABLE 3

Kitchen and Bath Surfaces

Comparative Bonding Test Results

Example	Surface	Organosilane Quat Test Results		
		Without H ₂ O ₂	With H ₂ O ₂	H ₂ O ₂ Alone
14	Baked Enamel	Poor	Good	Poor
15	Smooth Ceramic Tile	Fair	Good	Poor
16	Textured Ceramic Tile	Fair	Good	Poor
17	Chemically Frosted Glass	Good	Excellent	Poor
18	Formica Counter Top	Poor	Good	Poor
19	Corian (by DuPont)	Fair	Good	Poor
20	Marble Tile	Fair-Good	Good-Excellent	Poor
21	Granite Tile	Poor	Good	Poor
22	Glass	Good	Excellent	Poor

These results again demonstrate the invention on other kitchen and bath surfaces. In all examples the performance of the inventive organosilane quat composition with hydrogen peroxide demonstrated improved bonding or durability of the coating compared to the composition without hydrogen peroxide or to hydrogen peroxide treatment alone. With reference to Table 3, a synergism was observed between 3-(trimethoxysilyl)propyldimethyloctadecyl ammonium chloride and hydrogen peroxide on all surfaces. In all Examples 14-22, H_2O_2 , alone resulted in "Poor" bonding or durability. In Examples 14-16, 18-19 and 21, the organosilane quat, alone, also resulted in "Poor" or "Fair" bonding or durability. However, synergistically, the combination of the organosilane quat and hydrogen peroxide gave "Good" to "Excellent" results. Even in the case of Examples 17, 20 and 22, "Excellent" results were achieved for the inventive combination where H_2O_2 was not expected to contribute a beneficial result.

It was also noted during the test that prior to scrubbing the treated surface with MIRACLE SCRUB, all Examples gave an Excellent rating on the water drop test except for Example 19 which were rated as Good-Excellent on both surfaces. Also, all surfaces treated with hydrogen peroxide alone gave Poor results prior to scrubbing the treated surface with MIRACLE SCRUB.

EXAMPLES 23 - 35 - CLEANING OF EVERYDAY SURFACES

Test: To determine the ability of the composition (Example #1, Table #1, with hydrogen peroxide) to clean everyday spills and splatters off of previously untreated surfaces on which milk, coffee, tea, grape juice, soap scum, toothpaste residue and hard water minerals (as applicable to the surface) had been allowed to dry by evaporation for a period of at least 12 hours at an ambient temperature of $\pm 78^{\circ}\text{F}$. The composition was sprayed onto the soiled surfaces, allowed to penetrate the soil until the foam subsided and the surfaces were agitated and wiped dry using a soft paper towel.

Results: After wiping, the surfaces were examined both visually and with 3X magnification, and the paper towels were examined for stains, discoloration and residue typical of the soil that had been cleaned off the surface. The results are reported in Table 4.

TABLE 4

Example	Everyday Surface	Visual	3X Magnification	Paper Towel
23	Granite Counters and Tiles	Clean	No Residual	Soiled
24	Formica-type Counters	Clean	No Residual	Soiled
25	Glass Tables	Clean	No Residual	Soiled
26	Ceramic Counter, Wall and Floor Tiles	Clean	No Residual	Soiled
27	Polished Marble Tables and Counters	Clean	No Residual	Soiled
28	Exterior Windows and Patio Doors	Clean	No Residual	Soiled
29	Interior Windows, Mirrors and Patio Doors	Clean	No Residual	Soiled
30	Porcelain Sinks and Bathtubs	Clean	No Residual	Soiled
31	Sink, Bathtub and Shower Plumbing Fixtures	Clean	No Residual	Soiled
32	Acrylic, Fiberglass and Cultured Marble Sinks, Bathtubs and Shower Surrounds	Clean	No Residual	Soiled
33	Vehicular Windshields, Windows and Mirrors and Lights (Exterior)	Clean	No Residual	Soiled
34	Vehicular Windshields, Windows and Mirrors and Lights (Interior)	Clean	No Residual	Soiled
35	Vehicular Exteriors - Paint Finishes, Plastic and Metal Components	Clean	No Residual	Soiled

EXAMPLES 36 - 48- IMPARTING WATER AND SOIL REPELLENCY TO EVERYDAY SURFACES

Test: To determine the ability of the composition (Example #1, Table #1, with hydrogen peroxide) to clean everyday spills and splatters off of previously untreated surfaces and simultaneously impart a functional water and soil repellent finish to reduce the adhesion and buildup of subsequent exposure to everyday spills and splatters. Previously untreated surfaces, on which milk, coffee, tea, grape juice,

soap scum, toothpaste residue and hard water minerals (as applicable to the surface) had been allowed to dry by evaporation for a period of at least 12 hours at an ambient temperature of $\pm 78^{\circ}\text{F}$, were sprayed with the composition. The composition was allowed to penetrate the soil until

5 the foam subsided and the soiled areas were agitated and wiped dry with a soft paper towel. Additional quantities of milk, coffee, tea, grape juice, soap scum, diluted toothpaste and mineral-laden water were then spilled and splattered on the surface and allowed to evaporate for 12 hours at ambient temperature of 78°F . To assess the practical soil repellency

10 imparted by the composition, approximately one-half of the now treated and resoiled areas were dry-wiped with a soft paper towel to remove the spills and splatters. To assess the practical water and soil repellency imparted by the composition, the other half of the now treated and resoiled areas were sprayed with tap water and wiped dry.

15 Results: After dry-wiping, the surfaces were examined both visually and with 3X magnification for residual soil. After spraying with tap water and wiping until dry, the surfaces were examined both visually and with 3X magnification for residual soil. All paper towels showed various degrees of stain, discoloration and residue typical of the soil that

20 had been cleaned off the surface. The results are reported in Table 5.

TABLE 5

Example	Everyday Surface	Dry Wipe Visual 3X Magnification	Wet Wipe Visual 3X Magnification
36	Granite Counters and Tiles	Clean-No Residual	Clean-No Residual
37	Formica-type Counters	Clean-Very Slight Residual	Clean-No Residual
38	Glass Tables	Clean-No Residual	Clean-No Residual
39	Ceramic Counter, Wall and Floor Tiles	Clean-No Residual	Clean-No Residual
40	Polished Marbles Tables and Counters	Clean-Very Slight Residual	Clean -No Residual
41	Exterior Windows and Patio Doors	Clean-No Residual	Clean-No Residual
42	Interior Windows, Mirrors and Patio Doors	Clean-No Residual	Clean-No Residual
43	Porcelain Sinks and Bathtubs	Clean-No Residual	Clean-No Residual
44	Sink, Bathtub and Shower Plumbing Fixtures	Clean-No Residual	Clean-No Residual
45	Acrylic, Fiberglass and Cultured Marble Sinks, Bathtubs and Shower Surrounds	Clean-Slight Residual	Clean-No Residual
46	Vehicular Windshields, Windows, Mirrors and Lights (Exterior)	Clean-No Residual	Clean-No Residual
47	Vehicular Windshields, Windows, Mirrors and Lights (Interior)	Clean-No Residual	Clean-No Residual
48	Vehicular Exteriors - Paint Finishes, Plastic and Metal Components	Clean-Slight Residual	Clean-No Residual

**EXAMPLES 49 - 61 - PERFORMANCE AND DURABILITY OF
WATER AND SOIL REPELLENT COMPOSITIONS**

Test: To determine the performance and durability of the
water and soil repellent finish imparted by the composition (Example #1,

Table #1, with hydrogen peroxide) when used to clean everyday spills

and splatters off of treated surfaces, the surfaces were soiled with milk, coffee, tea, grape juice, soap scum, toothpaste residue and hard water minerals (as applicable to the surface being tested) and these elements were allowed to dry by evaporation for a period of at least 12 hours at an ambient temperature $\pm 78^{\circ}\text{F}$. The soiled areas were then sprayed with tap water and wiped with a soft paper towel. This resoiling and cleaning with tap water cycle was repeated 15 times. When measuring the water repellency of everyday, installed surfaces, it is impractical to conduct water drop or droplet tests with distilled water. To replicate real world conditions, spray tests are conducted using ordinary spray bottles filled with municipally furnished tap water typical of the region in which the tests are conducted.

Results: The soiled areas were easily cleaned with minimal wiping effort. Examination of the paper towels revealed residue and discoloration typical of the elements that had been removed. Visual examination of the surfaces after cleaning revealed no soil residue. After cleaning, tap water was sprayed on the treated surfaces and the hydrophobicity of the surfaces was rated high (75° or more), medium (50° to 75°) or low (25° to 50°) as follows in Table 6.

TABLE 6

Example	Everyday Surface	Cycle 3	Cycle 6	Cycle 9	Cycle 12	Cycle 15
49	Granite Counters and Tiles	High	High	Medium	Medium	Medium
40	Formica-type Counters	High	Medium	Medium	Low	Low
51	Glass Tables	High	High	High	High	High
52	Ceramic Counter, Wall and Floor Tiles	High	High	High	High	High
53	Polished Marble Tables and Counters	High	High	Medium	Medium	Low
54	Exterior Windows and Patio Doors	High	High	High	High	High
55	Interior Windows, Mirrors and Patio Doors	High	High	High	High	High
56	Porcelain Sinks and Bathtubs	High	High	High	High	High
57	Sink, Bathtub and Shower Plumbing Fixtures	High	High	High	Medium	Medium
58	Acrylic, Fiberglass and Cultured Marble Sinks, Tubs and Surrounds	High	High	Medium	Medium	Low
59	Vehicular Windshields, Windows, Mirrors and Lights (Exterior)	High	High	High	High	High
60	Vehicular Windshields, Windows, Mirrors and Lights (Interior)	High	High	High	High	High
61	Vehicular Exteriors - Pain Finishes, Plastic and Metal Components	High	High	Medium	Medium	Medium

**EXAMPLE 62 - THE PREPARATION AND TESTING OF
A CONCENTRATED CLEANING AND
MULTIFUNCTIONAL COATING COMPOSITION**

55 gms of deionized water was weighed into a 250 ml
beaker employing a magnetic stirring bar for mixing. 18 gms of

isopropanol were added and mixed. The solution was adjusted to pH = 3 with 1 N. hydrochloric acid. 4.2 gms of a 72% solution of 3-(trimethoxysilyl)propyldimethyloctadecyl ammonium chloride solution in alcohol (Dow Corning 9-6346) were added with stirring followed by
5 22.9 gms of 35% hydrogen peroxide. The final composition contained 3.0% 3-(trimethoxysilyl)propyldimethyloctadecyl ammonium chloride and 8.0% hydrogen peroxide and had a pH of 3.

The composition was applied to a glass surface and the Bonding Test completed according to the defined protocol. The result
10 after the MIRACLE SCRUB was "Excellent". Prior to scrubbing with MIRACLE SCRUB, the surface also showed an "Excellent" drop repellency result.

**EXAMPLE 63 - ANTIMICROBIAL EFFICACY OF CATIONIC
ORGANOSILANE QUATERNARY AMMONIUM
15 COMPOUND/HYDROGEN PEROXIDE COMPOSITION
ON TILE SURFACES**

These tests determine the antimicrobial efficacy of the compositions of this invention on everyday surfaces contaminated with a mixture of microbiological organisms consisting of *Escherichia coli*,
20 *Staphylococcus aureus* and *Bacillus subtilis*. The composition tested was Example 1 (Table 1) consisting of 3-(trimethoxysilyl)propyldimethyloctadecyl ammonium chloride with hydrogen peroxide. The following Materials and Methods were employed to conduct the tests.

Materials

1. Microorganisms:
 - a. *Escherichia coli* (ATCC 11229)
 - b. *Staphylococcus aureus* (ATCC 25923)
 - c. *Bacillus subtilis* (ATCC 6363)
2. Media:
 - a. Blood agar plate (BAP); Tryptic Soy Agar amended with 5% defibrinated sheep blood
 - b. Butterfield's Buffer (BFB)
3. McFarland Turbidity Standard (VWR)
4. Sterile Supplies:
 - a. Sterile swabs
 - b. Micropipette and sterile tips
5. Equipment:
 - a. Incubator
 - b. Spiral Biotech Autoplate 4000
6. Paper towels
7. Tile surfaces
8. Example 1 composition (cationic organosilane quaternary ammonium compound with hydrogen peroxide) for testing

Methods

1. The Blood Agar Plates (BAP) were inoculated with fresh cultures of the test organisms to obtain confluent growth and incubated at 35°C for 24 hours. A suspension of each organism was prepared by transferring colonies of each organism into separate 10 ml tubes of Butterfield's buffer. The suspensions were adjusted to a

turbidity of 0.5 McFarland units, corresponding to approximately 1×10^6 cells.

2. A 1:100 dilution of each suspension was performed by aliquoting 0.1 ml of suspension into 10 ml BFB, creating organism
5 suspensions.

3. To ensure no background microbial contamination was present, all test surfaces were cleaned with isopropyl alcohol and allowed to air dry.

4. Duplicate 100cm^2 areas were measured and
10 designated on the test surfaces for each organism and labeled using masking tape.

5. Test areas were sampled with BFB pre-moistened sterile swabs to confirm test areas were negative for bacterial growth.

6. Using a micropipette and sterile tips, 0.1 ml of the
15 organism suspensions were applied independently to the corresponding sections of the test surfaces and evenly distributed using a flame sterilized glass rod to cover the entire 100cm^2 .

7. Each surface was then treated with the composition of Example 1 by spraying it onto the surface to be cleaned and wiping
20 with paper towel or soft cloth until crystal clear and sparkling."

8. Test areas were sampled with BFB pre-moistened sterile swabs.

9. The organism suspensions and test sample swabs were vortexed to ensure homogeneity and cultured onto BAP via the Spiral Biotech Autoplate 4000.

10. The cultured samples were incubated at 35°C for 48 hours. Plates were examined and organisms enumerated.

Results

The analytical data for the tile surface are shown below in Table 7.

Table 7

Organism	Negative Surface Control	Organism Inoculum (0.1 ml of Organism Suspension)	Treated Surfaces
	CFU/100cm ²	CFU/ml	CFU/100cm ²
<i>E. coli</i>	< 10	2.88 x 10 ²	< 10
<i>E. coli</i>	< 10	2.88 x 10 ²	< 10
<i>S. aureus</i>	< 10	3.44 x 10 ²	< 10
<i>S. aureus</i>	< 10	3.44 x 10 ²	< 10
<i>B. subtilis</i>	< 10	7.27 x 10 ²	< 10
<i>B. subtilis</i>	< 10	7.27 x 10 ²	< 10

The negative control data demonstrates that all surfaces were free of bacteria prior to the experiment. The results indicate that the composition is an effective bactericide against *Bacillus subtilis*, *Staphylococcus aureus*, and *Escherichia Coli*.

Those of ordinary skill in the art realize that the descriptions, procedures, methods and compositions presented above can be revised

or modified without deviating from the scope of the described embodiments, and such do not depart from the scope of the invention.

WHAT IS CLAIMED IS: